Amendments to the Specification:

Please replace paragraph 2 on page 4 with the following amended paragraph:

In the present invention, basic copper carbonate which is a material for easily soluble copper oxide may be commercially available. Alternatively, it may be obtained by mixing an aqueous solution of copper chloride, copper sulfate or copper nitrate with an aqueous carbonate solution of, for example, alkaline metal, alkaline earth metal or ammonia_(NH4)_and then reacting the aqueous solutions with each other while heating them. In this instance, mixing between the aqueous carbonate solution and the aqueous solution of copper chloride, copper sulfate or copper nitrate may be carried out by charging the carbonate in the form of a solid into the aqueous solution of copper chloride, copper sulfate or copper nitrate to dissolve the former in the latter or charging copper chloride, copper sulfate or copper nitrate in the form of a solid into the aqueous carbonate solution to dissolve the former in the latter. The copper electroplating material of the present invention is soluble or easily dissolved in an electrolyte, to thereby be suitable for use as a copper plating material fed to the electrolyte in which an insoluble anode and a plated object acting as a cathode are arranged. In this instance, when the copper plating material contains such impurities as described above in a large amount, an article plated is deteriorated in quality. Thus, when the basic copper carbonate contains impurities such as alkaline metal (N a, or K), alkaline earth metal (Mg, or Ca) and/or an anion (Ct, or 8042-) or the like in a large amount, the easily soluble copper oxide is preferably washed with water.

Please replace the last paragraph on page 7 bridging, page 8, with the following amended paragraph:

The basic copper carbonate thus formed is then deposited or precipitated in the form of a powder. Then, a valve ~~2~~ 12 is rendered open to draw out the thus-precipitated slurry therethrough to a centrifugal separator 2, in which the slurry is subjected to centrifuging to separate a solid matter of the slurry from a mother liquor thereof. Then, the solid matter is placed in a drier 3, to thereby be dried therein, resulting in basic copper carbonate being obtained in the form of a powder.

Please replace the first paragraph on page 15 with the following amended paragraph:

Copper electroplating was executed under the following conditions while using copper oxide having a chlorine concentration_(CI concentration) of about 20 ppm as a copper supply.

Please replace the second paragraph on page 17 with the following amended paragraph:

A reaction temperature of the mixed solution_(temperature of the mixed solution) in the reaction tank 9 which is 70°C or below would permit a chlorine concentration of the basic copper carbonate to be reduced when the reaction time is increased. However, a reduction in chlorine concentration to a level below a reference level predetermined by the present invention is not attained even by the reaction extending over 8 hours, as will be noted from examples described below; thus, the reaction temperature at 70°C or below is not commercially accepted. On the contrary, the reaction temperature at 75°C leads to a satisfactory reduction in chlorine concentration by the reaction for, for example, 1.5 hours or more. The chlorine concentration tends to be decreased with an increase in reaction temperature when the reaction time is rendered the same. However, in the illustrated embodiment, the reaction temperature above 95°C causes an increase in chlorine concentration as will be noted from the examples described below. Also, even when the reaction temperature is set at a target value, it is unavoidable that the temperature is actually somewhat varied. Thus, in the illustrated embodiment, the reaction temperature, the target value is necessarily set to be within a range of between 75°C and 90°C.

Please replace the last paragraph beginning on page 18, bridging page 19, with the following amended paragraph:

More specifically, when a target level of the pH at 75°C is set to be 8.0, the feed ratio is 2.0; whereas when the target level at 95°C is set at 8.0, the feed ratio is, for example, 1.2. This is not due to dependence of the pH on a temperature because the solution having a pH of 8.0 at 100°C still has a pH (detected pH) of8.0 when the temperature is reduced to a level of 75°C. Thus, it is supposed that even when the aqueous solution has an apparent pH_(detected pH) of 8.0 at about 95°C, the actual pH is different from the apparent pH. This means that when the actual pH is 8.0, the apparent pH is out of 8.0. This would cause the feed ratio at 95°C to be considerably different from that at 75°C, resulting in the aqueous cupric chloride solution being excessively fed as compared with the aqueous sodium carbonate solution, leading to an increase in CI concentration.

Please replace the last paragraph on page 23, bridging page 24, with the following amended paragraph:

Basic copper carbonate was obtained by substantially repeating the procedure of Example 5-1 except that the reaction temperature was set to be each of 75°C, 80°C, 90°C, 95°C and

100°C. Then, a concentration of Cl in each of the basic copper carbonates thus obtained was measured. Results of the measurement were as shown in Table 7. (Results in connection with 75°C, 80°C and 90°C were described above.) Also, a feed ratio of the aqueous sodium carbonate which constitutes an alkali side to the aqueous cupric chloride solution which constitutes an acid side_(fed amount of an aqueous sodium carbonate solution divided by fed amount of an aqueous cupric chloride solution) was as shown in Table 7.